FORMULA WRITING & BOND TYPES

Part 1: Formula Writing Practice

1.

2.

3.

4.

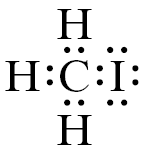
5.

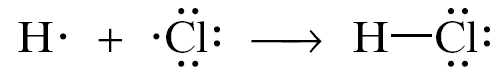
6.

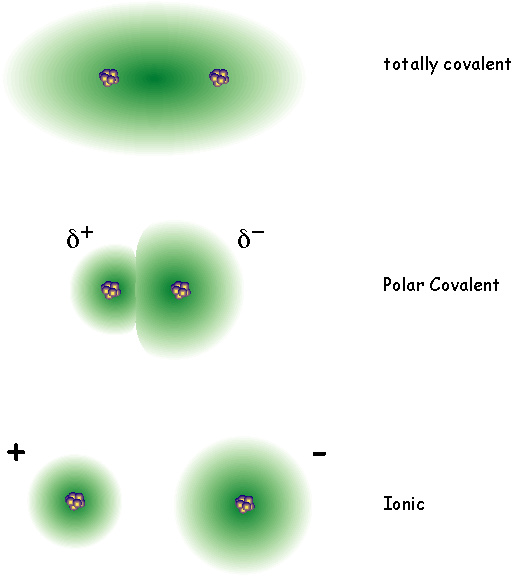
Part 2: Video Introduction

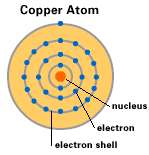
1. Electronegativity B. Oxidation Number C. Valence Electron(s) D. Lewis Dot Structure E. Lone Pair

\_\_\_\_ Na: +1

\_\_\_\_ 

\_\_\_\_ 

\_\_\_\_

\_\_\_\_ [](http://www.bing.com/images/search?q=valence+electrons&view=detail&id=628B66F71FDB125130F8AC389CA842C6A48D85FE&first=61&FORM=IDFRIR)

Part 3: Video

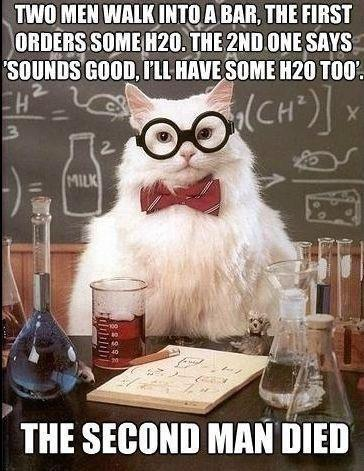
Generate at least 2 questions that you have from watching this video. Be prepared to share in class.

Part 4: Practice-On-Own

1. Using your knowledge of valence electrons and ionic charges, complete one of the following assignments. Be thorough and use these concepts accurately and clearly. Be creative and have fun.
2. Write a one paragraph love letter to calcium from oxygen. Keep it clean.
3. Acting as krypton’s parent, write a one paragraph note to krypton explaining why s/he is not allowed to date nickel.
4. Write a one paragraph news release how hydrogen came to chlorine’s aid and shared valence electrons with him/her.
5. Draw the Lewis Dot Structure for:

Hints: H will always be on the outside, C is typically a central atom, and most atoms need 8 valence electrons around them! Good luck.

CCl4 CH3CHCH2

1.  Given the chemistry cartoons below and your knowledge of the topic, write an explanation for these cartoons as if explaining to someone who does not know any chemistry!

